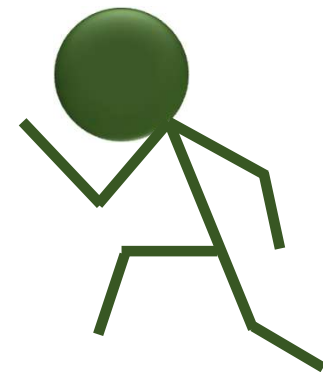


Assumption in kinetic theory of gases

- 1) Gases molecules are very small and occupy negligible volume compare to the total volume occupy by the gas.
- 2) Gases molecules are constantly moving in all direction and keep on colliding between their self and container wall.
- 3) Gas molecules undergo perfectly elastic collisions.
- 4) Only energy between the gas molecules is Kinetic Energy
- 5) Molecular speed is proportional to temperature.



Gas Molecule (No attraction force)